



# Reynolds High School

*Building Relationships for Academic Success*

## Chemistry 2024-2025

Teacher: Brigitte Jensen

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Credit: 1.0 Science Credit (0.5 per Semester)

**Required Materials:** Paper/notebook, pencil, computer supplemental materials provided

**Course Description:** Chemistry is the study of the structure and composition of matter that makes up living things and their environment. Students will also explore changes of matter and the mechanisms by which the changes occur.

### Class Expectations (i.e. POWER):



- 1. PREPARED and PUNCTUAL:** Come to class every day, on time and ready to learn!
- 2. ORGANIZED:** Turn in assignments and homework on time. Get help from me before an assignment is due. Use your planner.
- 3. WRITE:** Listen and speak to others but do your own best work. Copied work will not receive credit.
- 4. ENGAGED:** Most of the learning takes place in class, so PARTICIPATE 100%.  
**No electronics** (i.e. cell phones, music) can be used in class unless approved by teacher for academic purposes. Student using phones during class may result in the phone being taken and held per district policy.
- 5. RESPECT:** Show respect to everyone. Be respectful of your classmates' space and belongings and of my room. Be safe especially during labs. Follow adult directions the 1<sup>st</sup> time.
  - No food (drinks are okay in a sealed container).
  - Obey all laboratory safety rules and clean up after yourself and others.
  - Bathroom: 10/10 rule – not during first or last 10 minutes of class. Planners are your pass.

**Consequences:** If you do not meet the expectations of this class, you may run into the following consequences: verbal reminder, time out in the hall, call home, detention, meeting with parent and student, or a referral.

### Grading Policy- Overview

**Grading Policy:** Your grade in this class will be based upon your ability to **perform** skills, **describe** concepts, and **demonstrate** understanding of the **Learning Targets** for each unit. Learning Targets are based on the State and National standards for this course. Points earned are **weighted** in the following way:

- Formative Assessments (Classwork/ Homework/Labs)** are worth **30%** of your grade. **This** includes classwork, labs, homework as well as projects.
- Summative Assessments (Tests/Labs/Finals)** are worth **70%** of your grade: unit tests, lab practical, and final exams.

Highly Proficient

Proficient

Not Proficient

**A** 90 – 100%

**B** 80 – 89%

**C** 70 – 79%

**D** 60 – 69%

**F** < 60%

## Attendance

1. Be in your seat and ready when the bell rings.
2. *If you miss class check the “month ahead calendar” in Schoology for missed assignments.*
3. Labs should be done immediately upon return as materials will no longer be available after 1 week.
4. You are responsible for making up assignments from excused absences within one week.
5. Missed labs must be completed before or after school or during lunch.
6. Late work (not excused absence) will be accepted for reduced credit. Deadline for late work is the following unit test.

## Tentative Schedule

# weeks	Unit	Topics
4	Tools of Chemistry	<ul style="list-style-type: none"><li>▪ Chemistry and Matter</li><li>▪ How Scientists Study</li><li>▪ Scientific Notation</li><li>▪ Units of Measurement</li><li>▪ Making Measurements in the Lab</li></ul>
4	Atoms and the Periodic Table	<ul style="list-style-type: none"><li>▪ Atomic Theory</li><li>▪ Elements</li><li>▪ Mendeleev</li><li>▪ Organization of the Periodic Table</li><li>▪ Structure of the Atom</li><li>▪ Isotopes</li><li>▪ Electrons</li><li>▪ Periodic Table Trends</li></ul>
4	Ions and Ionic Compounds	<ul style="list-style-type: none"><li>▪ Octet Rule</li><li>▪ Ions - monatomic and polyatomic</li><li>▪ Ionic Bonding</li><li>▪ Writing Formulas</li><li>▪ Naming Compounds</li><li>▪ Properties of Ionic Compounds</li><li>▪ Acids and Bases</li><li>▪ Acid Rain</li></ul>
5	Covalent Bonds and Molecular Compounds	<ul style="list-style-type: none"><li>▪ Sharing Electrons</li><li>▪ Single, Double, Triple bonds</li><li>▪ Electronegativity and Bond Polarity</li><li>▪ Properties of Molecular Compounds</li><li>▪ Naming Binary Molecular Compounds</li><li>▪ The Love Canal</li></ul>
5	Chemical Reactions	<ul style="list-style-type: none"><li>▪ Law of Conservation</li><li>▪ Writing and Balancing Chemical Equations</li><li>▪ Quantitative Relationships</li><li>▪ Synthesis, Decomposition</li><li>▪ Single, Double Replacement</li><li>▪ Oxidation and Reduction</li><li>▪ Air Pollution</li></ul>
5	Quantities in Chemical Reactions	<ul style="list-style-type: none"><li>▪ The Mole</li><li>▪ Atomic and Molecular Mass</li><li>▪ Mole-Mass Conversions</li><li>▪ Mole-Mole Relationships</li><li>▪ Limiting Reagent and Percent Yield</li><li>▪ Global Climate Change</li></ul>
5	Solutions	<ul style="list-style-type: none"><li>▪ Types of Solutions</li><li>▪ Solubility</li><li>▪ Henry's Law</li><li>▪ Concentration</li><li>▪ Dilutions</li><li>▪ Ion Concentration</li><li>▪ Lead Pollution</li></ul>
2	Environmental Project	

## RHS Science Department - Laboratory Safety Agreement

I, \_\_\_\_\_ agree to abide by the

Student Name

following safety rules whenever working in the science laboratories.

I will:

1. Use the science laboratory for authorized work only.
2. Remove contact lenses and wear safety goggles when instructed.
3. Study the laboratory investigation before coming to class, if possible. (If in doubt about any procedure, I will ask the teacher.)
4. Know how to use safety equipment and the location of the eyewash station, safety shower and fire blanket.
5. In case of fire, alert the teacher and leave the laboratory.
6. Carefully check for the presence of ignition sources (open flames, etc.) before using flammable materials such as alcohol.
7. Report any accident, injury, spill, unsafe procedure or broken glass to the teacher at once.
8. Never taste, touch or smell any substance unless specifically directed by the teacher to do so.
9. Handle chemicals carefully, check the label of every bottle or jar before removing the contents, and never return chemicals to reagent containers.
10. When heating a substance in a test tube, point the mouth of the tube away from all persons.
11. Use proper equipment to handle hot glassware.
12. Tie back long hair, remove dangling jewelry, roll up loose sleeves, and tuck in loose clothing.
13. At the end of the lab, clean the work area, wash and store all materials and equipment, and turn off all water, gas, and electrical appliances.

We have read the laboratory safety agreement, the course syllabus, and acknowledge the content.

\_\_\_\_\_  
(Student Signature)

\_\_\_\_\_  
(Parent Signature)

\_\_\_\_\_  
date

\_\_\_\_\_  
date